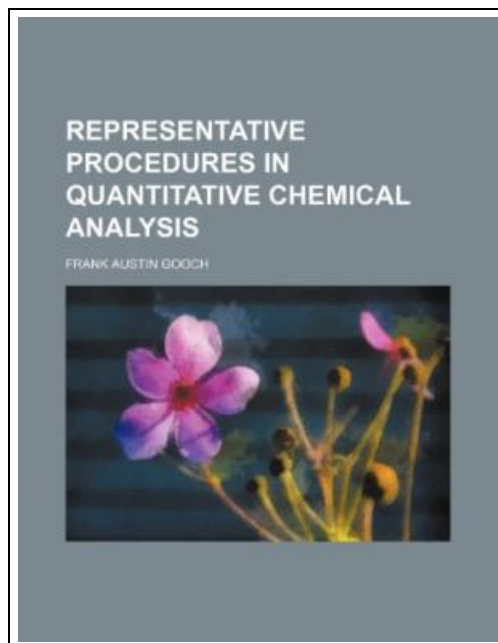


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Rarebooksclub.com, United States, 2012. Paperback. Book Condition: New. 246 x 189 mm. Language: English . Brand New Book ***** Print on Demand *****.This historic book may have numerous typos and missing text. Purchasers can download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1916 Excerpt: .indicator. The Standard Solution of Hydrochloric Acid.--To make a solution of hydrochloric acid of approximately standard strength it is only necessary to dilute the concentrated acid suitably with water. If pure concentrated hydrochloric acid is diluted until the specific gravity of the solution is about 1.020 a nearly normal solution is obtained, containing approximately 1 gm. mol. or 36.468 gm. to the liter. From such a solution more dilute solutions of strength suitable for use in titrimetric processes, N/5 or N/10, may be made by properly diluting measured portions with water. The strength of such solutions is then determined exactly by precipitating and weighing in the manner previously described for the gravimetric determination of the chloride ion, the silver chloride obtained from an exactly measured portion of the solution. Experimental Process: Standardization of Hydrochloric Acid.--From an approximately N/10 solution of hydrochloric acid, made by diluting 10 cm.³ of the concentrated acid of sp. gr. 1.02 to a liter, draw from a burette a 100-cm.³ portion, dilute it with water to a volume of 300 cm.³, add a few drops of nitric acid and then (gradually and with constant stirring to coagulate the precipitate) a 5 per cent solution of silver nitrate, until no further precipitation takes place. Heat the liquid to boiling and set the whole aside (in the dark) to cool. When completely settled, the precipitate is filtered off on asbestos in a perforated crucible, previously weighed, washed thoroughly with water containing a few drops...



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